Institute of Thermodynamics and Thermal Process Engineering Prof. Dr.-Ing. J. Gross



Classical Density Functional Theroy for Liquid-Fluid Interfaces using modified iSAFT and the heterosegmented GC-PC-SAFT Equation of State

Jonas Mairhofer¹ and Joachim Gross 1

¹Institute of Thermodynamics and Thermal Process Engineering, University of Stuttgart

Objective

- determine density profiles on a segment level to study the molecular structure at interfaces in detail
- good agreement with experimental data including associating components
 consistency with bulk equation of state (here: heterosegment

Results

Surface Tensions of non-associating components
35
30
25
× Experiment
DFT

$GC-PC-SAFT^{1}$)

Classical Density Functional Theory (DFT)

 Grand potential of an inhomogeneous system with a planar interface and no external potential

$$\Omega[\rho_k(z)] = F[\rho_k(z)] - \sum_{j=1}^{NS} \mu_{jc} \int \rho_j(z) dz \qquad (1)$$

• At thermodynamic equilibrium for given T, V, μ_{jc}

$$O = \frac{\delta \Omega[\rho_k(z)]}{\delta \rho_i(z)} = \frac{\delta F[\rho_k(z)]}{\delta \rho_i(z)} - \mu_{ic}^{eq}$$
(2)

 The equilibrium density profiles are obtained by iterating the following equation

$$\rho_{is}(z) = \exp\left(\beta\mu_{ic}^{Eq} + D_{is}(z)\right) I_{1,is}(z)I_{2,is}(z)$$
(3)

Helmholtz Energy Functional

Residual Helmholtz energy functional



Figure : Comparison of DFT to experimental data of 1-propene⁷, butane⁷, hexane⁷, octane⁷ and dodecane⁷ (from left to right).

Surface Tensions of associating components



 $F^{\rm res}[\rho_k(z)] = F^{\rm hs}[\rho_k(z)] + F^{\rm hc}[\rho_k(z)] + F^{\rm disp}[\rho_k(z)] + F^{\rm asso}[\rho_k(z)] \quad (4)$

Hard sphere contribution: modified fundamental measure theory²³
 Chain formation: modified iSAFT (Jain, Dominik and Chapman⁴)
 Dispersive contribution: weighted density approach (Sauer)

$$\beta F^{\text{disp}}[\rho_k(z)] = \int \sum_{j=1}^{NS} \rho_j(z) \bar{a}_j^{\text{disp,GC-PC-SAFT}}(\bar{\rho}_k(z)) dz \qquad (5)$$
$$\bar{\rho}_i(\mathbf{r}) = \frac{3}{4\pi \psi_i^3} \int \rho_i(\mathbf{r'}) \Theta(\psi_i - |\mathbf{r'} - \mathbf{r}|)(d) \mathbf{r'} \qquad (6)$$

• The value of ψ_i was adjusted for a DFT consistent with PC-SAFT REF and is used unchanged here

Associative contribution: iSAFT (Bymaster and Chapman⁵)

$$\beta F^{\text{Asso}}[\rho_k(z)] = \int \sum_{j=1}^{NS} \rho_j(z) \sum_{A \in \Gamma^{(j)}} \left[\ln \chi^j_A(z) - \frac{\chi^j_A(z)}{2} + \frac{1}{2} \right] dz \quad (7)$$

• Equation for χ^i_A was modified to be consistent with bulk equation of state

$$\chi_{A}^{i}(z) = \frac{1}{1 + \frac{1}{2} \sum_{j=1}^{NS} \kappa_{ij} \sigma_{ij}^{2} \int_{z-\sigma_{ij}}^{z+\sigma_{ij}} \rho_{j}(z') \sum_{B \in \gamma^{(j)}} \chi_{B}^{i}(z') \{ [\exp(\beta \epsilon_{Ai,Bj}) - 1] y^{ij}(z,z') \} dz'$$
(8)



Figure : Comparison of DFT to experimental data of butylamine⁸, pentylamine⁸, 1-hexanol⁹ and tetradecanoic acid¹⁰ (from left to right).

Density Profiles at Vapor-Liquid Interface



Figure : Density Profiles of the single segments of butylamine at T = 300K.

Figure : Density Profiles of the single segments of 1-butanol at T = 360K.

Funding

The research leading to these results has received funding from the European Community's Seventh Framework Programme (FP7/2007-2013) under grant agreement n° [604271].

Numerical method

The resulting system of equations, eq. 3, is solved in parallel using an inexact Newton method from PETSc⁶.



Figure : Absolute residuals over time for the surface tension calculation of dodecane at T = 450 K using different numbers of processors.

References

[1] J. Gross, O. Spuhl, F. Tumakaka, and G. Sadowski *Ind. Eng. Chem. Res.*, vol. 42, no. 6, pp. 1266–1274, 2003.

[2] R. Roth, R. Evans, A. Lang, and G. Kahl *J. Phys. Condens. Matter*, vol. 14, no. 46, p. 12063, 2002.
[3] Y.-X. Yu and J. Wu *J. Chem. Phy.*, vol. 117, no. 22, pp. 10156–10164, 2002.

[4] S. Jain, A. Dominik, and W. G. Chapman *J. Chem. Phys.*, vol. 127, no. 24, p. 244904, 2007.

[5] A. Bymaster and W. G. Chapman *J. Phys. Chem. B*, vol. 114, no. 38, pp. 12298–12307, 2010.

[6] S. Balay, S. Abhyankar, M. F. Adams, J. Brown, P. Brune, K. Buschelman, L. Dalcin, V. Eijkhout, W. D. Gropp, D. Kaushik, M. G. Knepley, L. C. McInnes, K. Rupp, B. F. Smith, S. Zampini, H. Zhang, and H. Zhang, "PETSc Web page." http://www.mcs.anl.gov/petsc, 2016.

[7] P. J. Linstrom and W. G. Mallard, "NIST Chemistry WebBook," in NIST Standard Reference Database Number 69, National Institute of Standards and Technology. http://webbook.nist.gov (retrieved May 10, 2016).

[8] A. Vogel J. Chem. Soc., pp. 1825–1833, 1948.

[9] Y. V. Efremov Russ. J. Phys. Chem., vol. 40, no. 6, p. 667, 1966.

[10] K. Hunten and O. Maass J. Am. Chem. Soc., vol. 51, pp. 153–165, 1929.